

(19)



Europäisches Patentamt

European Patent Office

Office européen des brevets



(11)

EP 1 094 030 A2

(12)

EUROPEAN PATENT APPLICATION

(43) Date of publication:

25.04.2001 Bulletin 2001/17

(51) Int. Cl.⁷: **C01B 3/38**, **C01B 3/32**,

C01B 3/48, **B01J 8/04**

(21) Application number: **00122651.3**

(22) Date of filing: **18.10.2000**

(84) Designated Contracting States:

**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE**

Designated Extension States:

AL LT LV MK RO SI

(30) Priority: **20.10.1999 JP 29794599**

03.03.2000 JP 2000058483

(72) Inventor: **Maruko, Saburo**

Yamato-shi, Kanagawa (JP)

(74) Representative:

Fuchs Mehler Weiss & Fritzsche

Patentanwälte

Postfach 46 60

65036 Wiesbaden (DE)

(71) Applicant:

**NIPPON CHEMICAL PLANT CONSULTANT CO.,
LTD.**

Chiyoda-ku Tokyo (JP)

(54) **Reforming method and apparatus for hydrogen production**

(57) An auto-oxidation and internal heating type reforming method and apparatus for hydrogen production are disclosed for use in a process in which a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water vapor is fed into contact with a mass of a reforming catalyst to bring about a reforming reaction of the gaseous mixture to produce hydrogen, wherein a small amount of an oxidizing catalyst is admixed with the reforming catalyst in that mass; and a small amount of oxygen is admixed with the gaseous mixture, whereby a portion of the hydrocarbon or aliphatic alcohol is exothermally oxidized to generate a quantity of heat required to reform the gaseous mixture of the hydrocarbon or aliphatic alcohol with water vapor.

EP 1 094 030 A2

Description

Technical Field

[0001] The present invention relates to improvement in a method of and an apparatus for producing hydrogen by a reforming reaction. In particular, it relates to providing an auto-oxidation and internal heating type reforming method and apparatus for hydrogen production.

Background Art

[0002] So far in producing hydrogen, a reforming method and apparatus using a reforming reaction has advantageously been adopted in which a mixture of a hydrocarbon or an aliphatic alcohol with water vapor is fed onto and contacted with a reforming catalyst. In such methods and apparatus, it is customary that a quantity of heat needed for the reforming reaction is provided externally via wall surfaces of the reforming reactor.

[0003] In such conventional methods and apparatus as described, the necessary amount of heat for the reforming reaction is supplied externally via the wall surfaces of the reforming reactor with a sensible heat of a combustion gas burnt outside of the reforming reactor or a sensible heat of a heating medium. It has therefore been necessary that an external combustion furnace or burner unit for heat supply or thermocirculating unit for the heating medium or thermal catalyst be separately provided. As a result, the entire apparatus has had to be large in size and yet it has been difficult to obtain well satisfactory thermal efficiency. Also, in the case of hydrocarbons use, methane whose reforming reaction temperature is required for 700 to 750 °C is usually used. And if methane is heated externally to reach that temperature in the reforming reactor, the wall surfaces of the reforming reactor must be limited in temperature to 1000 °C or less, otherwise the material that makes up those wall surfaces tends to deteriorate, rendering the reforming reactor not usable over a long period of time and thus causing a difficulty in its workability.

Disclosure of Invention

[0004] The present invention is contrived to solve the foregoing problems and it is an object of the present invention to provide a reforming method for hydrogen production which makes a reforming reactor apparatus compact, which largely reduces energy cost needed for reforming, and which is capable of bringing development of NO_x, an environmental pollutant, almost or unlimitedly to zero. It is also an object of the present invention to provide a reforming apparatus for carrying out such a method.

[0005] In order to achieve the objects mentioned above there is provided in accordance with the present

invention in a first aspect thereof an auto-oxidation and internal heating type reforming method for hydrogen production for use in a process in which a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water vapor is fed into contact with a mass of a reforming catalyst to bring about a reforming reaction of the gaseous mixture therewith to produce hydrogen, wherein the method comprises the steps of: admixing a small amount of an oxidizing catalyst with the said reforming catalyst in the said mass; and admixing a small amount of oxygen with the gaseous mixture of the hydrocarbon or aliphatic alcohol with water vapor, whereby a portion of the hydrocarbon or aliphatic alcohol is exothermally oxidized to generate an amount of heat required to reform the gaseous mixture of the hydrocarbon or aliphatic alcohol with water vapor.

[0006] These and other features, objects and advantages of the present invention will become more readily apparent to those of ordinary skill in the art from the following detailed description of the preferred forms of embodiment thereof as illustrated in the various drawing Figures.

Brief Description of Drawings

[0007] In the accompanying drawings,

Fig. 1 is a cross sectional view in elevation illustrating a reforming apparatus for use in carrying out a method according to the present invention, and as implemented in a first form of embodiment thereof; Fig. 2 is a diagrammatic view illustrating the reforming apparatus in the first form of embodiment of the invention shown in Fig. 1 and its related peripheral equipment;

Fig. 3 is a cross sectional view in elevation illustrating a reforming apparatus for use in carrying out a method according to the present invention, and as implemented in a second form of embodiment thereof;

Fig. 4 is a cross sectional view taken along the line IV-IV as viewed from top in Fig. 3; and

Fig. 5 is a cross sectional view in elevation illustrating a reforming apparatus for use in carrying out a method according to the present invention, and as implemented in a third form of embodiment thereof.

Best Modes for Carrying Out the Invention

[0008] Research and investigations by the present inventor have revealed that feeding a bed or layer of a reforming reaction catalyst with a gaseous mixture a hydrocarbon or an aliphatic alcohol with water vapor, if the gaseous mixture has a small amount of oxygen added thereto and the catalytic layer has a small amount of an oxidizing catalyst added thereto, permits a portion of the hydrocarbon or aliphatic alcohol to be exothermally oxidized and to internally derive an

amount of heat as required to reform the hydrocarbon or aliphatic alcohol.

[0009] In this case, a small amount of oxygen may be added to the gaseous mixture of a hydrocarbon or aliphatic alcohol with water vapor either in advance or after the gaseous mixture on contacting with the reforming catalyst is in part reformed. Alternatively, a portion of the small amount of oxygen may be added to the gaseous mixture of a hydrocarbon or aliphatic alcohol with water vapor in advance, and a remainder the small amount of oxygen may be added after the resultant gaseous mixture on contacting with the reforming catalyst is part reformed.

[0010] By the way, a large difference exists between a hydrocarbon and an aliphatic alcohol in the amount of an oxidizing catalyst used, because of a difference in temperature between their respective reforming reactions as required.

[0011] Thus, a hydrocarbon in order for its reforming reaction to be accomplished requires a temperature of 700 to 750 °C. In the conventional external heating reformer, a hydrocarbon commences its reforming reaction when its gaseous mixture with water vapor is externally heated to a temperature around 500 °C, the reforming reaction coming to an end substantially when the temperature of the reforming reaction is raised up to 700 to 750 °C.

[0012] Since the reforming reaction is an endothermic reaction, it is important that an amount of heat commensurate with the heat absorbed be replenished enough so that the reaction temperature may not fall but be maintained in the reforming reaction catalyst.

[0013] According to the method of the present invention, an exothermic oxidizing reaction that is caused to occur inside of the catalyst bed or layer has been found to produce a sensible heat which is sufficient to maintain the temperature substantially in a range between 700 °C and 800 °C and thus is capable of maintaining a temperature as required for the reforming reaction. To wit, if low temperature hydrogen produced by the reforming reaction brings about a temperature drop in the gaseous mixture of its raw material hydrocarbon with water vapor, oxygen that remains in the gaseous mixture will react with the hydrocarbon again in the presence of the oxidizing catalyst in the catalytic bed or layer to restore the gaseous mixture to a temperature of about 800 °C, thus continuing the reforming reaction repetitively. In this way, the method of the present invention, by giving rise to a state such as if a multi-stage fine catalytic combustion took place in a catalytic layer, is found to largely reduce even the amount of a reforming reaction catalyst and to make it possible to reduce the size of the apparatus as well. An amount of an oxidizing catalyst that ranges between 1 and 10 % of a reforming reaction catalyst (e. g., for methane, 3 % \pm 2 %, namely in a range of 1 to 5 %) has been found to be suitable for use. For an oxidizing catalyst use may be made of any catalyst that can withstand

a temperature in this range, although use is typically made of platinum, palladium or the like as distributed in the reforming catalyst.

[0014] An aliphatic alcohol is markedly lower in reforming temperature than a hydrocarbon. Not only does its reforming reaction go on at a temperature as low as 250 to 350 °C, but it absorbs less amount of heat for the reforming reaction, permitting the amount of oxygen being fed into a gaseous mixture of the aliphatic alcohol with water vapor to be considerably reduced. As a consequence, the reforming reactor can even more be reduced in size. For an aliphatic alcohol, an amount of an oxidizing catalyst that ranges between 1 and 5 % of a reforming reaction catalyst has been found to be suitable for use (e. g., for methanol, 2 % \pm 1 % is preferable).

[0015] For instance, if the reforming reaction catalyst has a space velocity SV of approximately 3,000 or so, an oxidizing catalyst having an SV of approximately 100,000 or so will achieve its purposes. Then, it is important that the oxidizing catalyst be dispersed and distributed enough evenly in the reforming reaction catalyst. Besides, to facilitate commencement of the reaction it has been found advantageous to dispose a thin layer of the oxidizing catalyst in front of the reforming reaction catalytic layer containing the oxidizing catalyst. Also, for concluding the reaction more quickly, it has been found that a thin layer of the oxidizing catalyst may advantageously be disposed behind the composite (mixed) reforming reaction catalytic layer.

[0016] A hydrocarbon or aliphatic alcohol fed into a reforming reaction catalytic bed or layer having an oxidizing catalyst admixed therein reacts with oxygen by the admixed oxidizing catalyst and is exothermally oxidized. A resultant gaseous mixture has its temperature elevated and, upon contacting with the reforming reaction catalyst, brings about a reforming reaction thereby to produce hydrogen. As mentioned before, the reforming reaction being an endothermic reaction causes the gaseous mixture to reduce in temperature, but the presence of the oxidizing catalyst in the reforming reaction catalytic bed or layer downstreams permits an unreacted portion of the hydrocarbon or aliphatic alcohol still further to react with such residual oxygen, thus generating heat that tends to maintain the temperature of the gaseous mixture substantially constant. In this way, the reforming reaction is allowed to continue in steps and in succession until the amount of oxygen in the gas fed may completely be consumed. In this case, dimensioning the thickness of the reforming reaction catalytic layer traversed by the gaseous mixture to be commensurate with the amount (rate) of feed of the gaseous mixture permits the reforming reaction to be concluded at the end of a way out of the reforming reaction catalytic bed or layer. From the standpoint of environmental control, it is desirable that the residual amount of oxygen be made completely nil by providing the exit end of the catalytic layer with a thin layer made up only of the oxidizing cat-

alyst.

[0017] The ranges of SV suitable for the reactions in the present invention have been found to lie in 1,500 to 8,000 for a hydrocarbon and 2,000 to 8,000 for an aliphatic alcohol.

[0018] In the case of hydrocarbons as a raw material in the present invention, methane (CH_4), ethane (C_2H_6), propane (C_3H_8), kerosene, gasoline and so forth can be used, and typically methane (CH_4) is used. Then, the reforming reaction generally proceeds at a temperature of 750 to 800 °C. For aliphatic alcohol for use in the present invention, use may be made of methanol, ethanol and so forth, and typically used is methanol. Then, the reforming reaction is carried out usually at a temperature of 250 to 350 °C.

[0019] The proportion of water vapor to a hydrocarbon: ($\text{H}_2\text{O}/\text{C}$) used typically ranges between 2.5 and 3.5. The proportion of water vapor to an aliphatic alcohol: ($\text{H}_2\text{O}/\text{C}$) used typically ranges between 2 and 3.

[0020] For a reforming reaction catalyst, use may be made of any one commonly used. Well often used may include $\text{NiS-SiO}_2 \cdot \text{Al}_2\text{O}_3$, $\text{WS}_2\text{-SiO}_2 \cdot \text{Al}_2\text{O}_3$ and $\text{NiS-WS}_2 \cdot \text{SiO}_2 \cdot \text{Al}_2\text{O}_3$.

[0021] Post-reforming gases include hydrogen, carbon dioxide, carbon monoxide and water vapor. To increase the hydrogen concentration, a shift reaction may be employed. As a shift catalyst, ordinarily Fe_2O_3 or Fe_3O_4 is used, but if the reaction temperature exceeds 700 °C, it is desirable to use Cr_2O_3 .

[0022] For an oxidizing catalyst, preference is given of platinum (Pt) or palladium (Pd) which is resistant to a deleterious change at an elevated temperature.

[0023] For oxygen to be added, preference is given of pure oxygen. For pure oxygen, use is made of oxygen produced from a water electrolytic cell or bath that concurrently produces hydrogen which is advantageously used here as admixed with hydrogen obtained from the reforming reactor apparatus in the present invention. Noting that a water electrolytic bath or cell has an efficiency nowadays increased up to 80 to 90 %, this combination is found to be industrially valuable. In economical view point, it might be possible to use air as a necessary oxygen resource. However, use of air reduces the concentration of hydrogen obtained by the reforming reaction. Thus, if air is used, it is desirable that an oxygen separation membrane be employed to increase the oxygen concentration in such air for feeding in the present invention.

[0024] Also, if air is used in reforming a gaseous mixture of a hydrocarbon with water vapor, such air should desirably have its operating pressure raised to 6 kg/cm^2 to 8 kg/cm^2 . Then, the reforming reaction temperature may also be raised from the aforementioned range of 700 to 750 °C further to a range of 800 to 900 °C.

[0025] This is because of the fact that if air is so used, it is necessary to later remove nitrogen (N_2) from a reaction gas in order to raise its hydrogen concentra-

tion. This is attained by cooling the reaction gas after both the reforming and shift reactions to a temperature around 55 °C and passing it through a hollow-fiber type separator membrane apparatus using, e. g., polysulfone. In this case, since the operating pressure for the reforming reaction to be raised in the range of 6 to 8 kg/cm^2 , the reforming reaction temperature inevitably should be raised from the normal range of 700 to 750 °C to a higher range of 800 to 900 °C.

[0026] If an oxygen-enriched air which oxygen concentration was raised by using the membrane-type separator apparatus, and an assumption is made that air is supplied under its pressure ranging between 8.4 kg/cm^2 and 9.0 kg/cm^2 , and that the yield of oxygen is to be 90 %, the resulted oxygen will have a concentration of 30 to 33 %. And if the yield of oxygen is 70 %, the resulted oxygen will have a concentration of 60 to 63 %. Operating conditions may be selected from these ranges depending on an oxygen concentration as required.

[0027] An explanation of the present invention is shown as below, specifically in respect of two forms of embodiment thereof. It should be noted, however, that these forms of embodiment are illustrative and not intended to limit the present invention thereto.

First Form of Embodiment

[0028] Mention is made of the first form of embodiment of the present invention with reference to Figs. 1 and 2. This form of embodiment of the invention is characterized in that a gaseous mixture of a hydrocarbon or aliphatic alcohol, with water vapor has oxygen added therein and admixed therewith in advance.

[0029] Mention is first made of a reforming reactor apparatus for practicing this form of the invention.

[0030] The reforming reactor apparatus includes a reforming reactor's outer cylinder 1 closed at bottom and a reforming reactor's inner cylinder 2 which define a space therewith, in which space a mixture 52 of a reforming reaction catalyst with a small amount of an oxidizing catalyst is loaded or packed in bed. The space has its lower end provided with a reforming reaction catalyst supporting lattice 4 as shown. A gaseous mixture having undergone a reforming reaction is allowed to pass through the lattice 4 and then pass a lower space to flow through an inner flow passage loaded with a high temperature shift catalyst 24, during which it is subjected to a shift reaction. The shift reaction is a reaction by which a mixture of carbon monoxide and water vapor is converted into hydrogen and carbon-dioxide gases and, it being an exothermic reaction, an amount of heat then produced acts to heat, via an outer wall 2 of the inner flow passage, an outer pack bed or catalytic layer of the reforming reaction catalyst and oxidizing catalyst mixture 52, namely, is fed into the latter. Using such a construction makes it possible, with heat generation anticipated, to reduce the amount of oxygen being added to the gaseous mixture of hydrocarbon or

aliphatic alcohol with water vapor.

[0031] A further explanation in detail is given below along the gas flow in respect of the first form of embodiment shown in Figs. 1 and 2.

[0032] A feed gas consisting of the gaseous mixture of hydrocarbon or aliphatic alcohol with water vapor and a small amount of oxygen added thereto and admixed therewith is fed through a reforming gaseous mixture pre-heater inlet 19 and, then flowing inside of a low temperature shift reactor jacket 8, is allowed to flow out through a reforming gaseous mixture pre-heater outlet 20. The feed gas then flows through a reforming gaseous mixture inlet 21 into a reforming reactor outer cylinder jacket 7. The feed gas while flowing through the low temperature shift reactor jacket 8 and the reforming reactor apparatus outer cylinder jacket 7 is pre-heated and then entering a reforming reactor cylinder or annular space 23 loaded with a reforming reaction catalyst and oxidizing catalyst mixture 52, undergoes a reforming reaction while in contact with the catalyst.

[0033] A gaseous mixture produced in the reforming reaction is passed thorough the reforming reaction catalyst supporting lattice 4, and then through a high temperature shift catalyst supporting lattice 5, and finally flowing into a high temperature shift catalyst loaded cylinder 24. With the layer of the reforming reaction catalyst and oxidizing catalyst mixture 52 in the reforming reaction cylinder or annular space 23 having a small amount of the oxidizing catalyst admixed therewith in dispersion, the gaseous mixture is raised in temperature with a heat generated by a multi-step oxidizing reaction that it undergoes when it passes through the layer, and thus permits the reforming reaction to continue in succession. A reaction gas having undergone the reforming reaction is reduced in temperature on exchanging heat in the lower space of the reforming reactor apparatus with the reforming gaseous mixture flowing through the outer cylinder jacket 7 of the reforming reactor apparatus and flows into the high temperature shift catalyst cylinder 24. A heat is generated by a shift reaction in the high temperature shift catalyst cylinder 24 and is transferred into the reforming reaction cylinder or annular space 23 through an inner wall thereof, thus serving to reduce the temperature drop there by the endothermic reaction. As a result, it is made possible to reduce the amount of oxygen being added to the gaseous of hydrocarbon or aliphatic alcohol with water vapor. Since the shift reaction is generally insufficient by a high temperature shift reaction, in practice a low temperature shift reaction is additionally carried out. Also, since a reaction gas leaving the high temperature shift reaction is raised significantly in temperature, it is desirable to cool the same in an heat-exchanging arrangement such that as shown in Fig. 2, water raised in temperature by exchanging heat in a water cooler C is made to flow through a gas cooler medium inlet 17 into a high temperature shift reaction cylinder outlet cooler 16. The cooled reaction gas is passed through a low

temperature shift catalyst layer or bed 25 to conclude the shift reaction and is taken out of the system through a reformed gas outlet 26. An amount of water vapor produced by heat exchange in the cooler 16 is taken out through a gas cooler medium outlet 18 for use in the reforming gaseous mixture by mixing it with hydrocarbon or aliphatic alcohol and oxygen in a mixer F as shown in Fig. 2.

10 Second Form of Embodiment

[0034] Mention is made of the second form of embodiment of the present invention with reference to Figs. 3 and 4. This form of embodiment of the invention is characterized in that oxygen is added to a gaseous mixture of a hydrogen containing reformable reactant or composition, here constituted by a hydrocarbon or aliphatic alcohol, with water vapor, after the gaseous mixture is in part reformed.

[0035] Mention is first made of a reforming reactor apparatus for practicing this form of the invention.

[0036] The reforming reactor apparatus includes a reforming reactor's outer cylinder 1 basically closed at bottom and more than one reforming reactor's inner cylinders 2 which are disposed spaced apart from each other in the outer cylinder 1. The outer and inner cylinders 1 and 2 define a space therewith, in which space a reforming reaction catalyst 51 is loaded or packed in bed. The packed or loaded reforming reaction catalyst 51 is supported by a reforming reaction catalyst support or carrier 4, e. g., in the form of a lattice, disposed in the region of the lower ends of the reforming reactor's inner cylinder 2. Loaded or packed in each of the reforming reactor's inner cylinder 2 are, to say in the order from upstream to downstream of the gas flow, a reforming reaction catalyst and oxidizing catalyst mixture 52 containing a small amount of an oxidizing catalyst, a high temperature shift catalyst 24 and a low temperature shift catalyst 25 which are supported by the reforming reaction catalyst support 4, a high temperature shift catalyst support 5 and a low temperature shift catalyst support 6, respectively.

[0037] Disposed underneath each of the reforming reactor's inner cylinders 2 is a mixer 38 for adding oxygen into the reformable gaseous mixture. Each mixer 38 has an opening to which oxygen blowing nozzle 39 sits opposite. While as illustrated in Fig. 3 the mixer 38 is shown as attached to an inner wall of the reforming reactor's inner cylinder 2 by means of a mixer fastening pin 55, it should be noted that the mixer may be of any other form and may be attached in any other way.

[0038] As illustrated in Fig. 3, two elongate reforming reactor's inner cylinders 2 are shown as disposed left and right, each extending vertically, and are also shown as having a reforming gas collecting tube 33 disposed above them as can also be seen from the top plan view of Fig. 4. The reforming gas collecting tube 33 so disposed and arranged has at its top center a branch

tube such that a reforming gas may be collected through it upwards.

[0039] A further explanation in detail is given below along the gas flow in respect of the second form of embodiment shown in Figs. 3 and 4.

[0040] A gaseous mixture of hydrocarbon or aliphatic alcohol with water vapor is introduced through reforming gaseous mixture inlet pipe 49 and allowed to flow through the space defined by the reforming reactor's outer cylinder 1 and the reforming reactor's inner cylinders 2.

[0041] This space is loaded with the reforming reaction catalyst 51.

[0042] A reforming reaction caused by the reforming reaction catalyst and oxidizing reaction catalyst mixture 52 loaded in the inner cylinders 2 results in a high temperature gas whose heat is transferred via walls of the inner cylinders 2 into the reforming reaction catalyst 51 where a portion of the reformable gaseous mixture of hydrocarbon or aliphatic alcohol with water vapor is reformed.

[0043] The reformable gaseous mixture in part reformed and passed through the reforming catalyst support 4, then enters into the mixers 38 where it is admixed with oxygen or air from the oxygen or air blowing nozzles 39. The admixture then passes through the reforming catalyst support 4 and flows into the space loaded with the reforming reaction catalyst and oxidizing reaction catalyst mixture 52 in the inner cylinders 2.

[0044] In this space, the reformable gaseous admixture in part reformed is raised in temperature by the action of the oxidizing reaction catalyst to a target reaction termination temperature and is then reformed by the action of the reforming reaction catalyst, thus terminating the reforming reaction.

[0045] The temperature the fully reformed gas is high and a heat held in that gas is transferred through the walls of the inner cylinders 2 to the reformable gaseous mixture, i. e., to the reforming reaction catalyst 51 in the space between the reforming reactor's outer cylinder 1 and the reforming reactor's inner cylinders 2, thus reforming the reformable gaseous mixture in part.

[0046] The fully reformed gas is passed through the reforming reaction catalyst and high temperature shift catalyst support 5 enters the region of the high temperature shift catalyst 24 where it undergoes a shift reaction.

[0047] Heat generated in this reaction is transferred as in the preceding step through the walls of the reforming reactor's inner cylinders 2 to the reformable gaseous mixture where it is utilized to raise the temperature of the reformable gaseous mixture.

[0048] The regions whose temperatures are lower and insufficient to cause the reformable gaseous mixture to be reformed may be simply loaded with ceramic or the like particles in lieu of the reforming reaction catalyst.

[0049] Next, the fully reformed gas whose heat is

transferred to the reformable gaseous mixture and thus which is lowered in temperature flows passed through the low temperature shift catalyst support 6 into the region loaded with the low temperature shift catalyst 25 where it completes the shift reaction.

[0050] The amount of heat that acts to lower any heat or sensible heat generated in the meantime is transferred via the walls of the reforming reactor's inner cylinders 2 to the reformable gaseous mixture for utilization to raise its temperature as well as have been mentioned.

Third Form of Embodiment

[0051] Mention is made of the second form of embodiment of the present invention with reference to Fig. 5. This form of embodiment of the invention is characterized in that a portion of a small amount of oxygen is added in advance to a gaseous mixture of a hydrocarbon or aliphatic alcohol with water vapor coming into contact with reforming catalyst, and the remainder of the small amount of oxygen is added to the gaseous mixture after the gaseous mixture is in part reformed.

[0052] Mention is first made of a reforming reactor apparatus for practicing this form of the invention.

[0053] The reforming reactor apparatus for use in this form of embodiment is shown in Fig. 5 as one shown in Figs. 3 and 4 and in part modified. In the modified arrangement, upstreams of the reforming catalyst bed 51 in the space formed between the reforming reactor's outer cylinder 1 and reforming reactor's inner cylinder 2, a region is provided loaded with a refining catalyst and oxidizing catalyst mixture 52' that contains a small amount of an oxidizing catalyst.

[0054] In the modified arrangement shown in Fig. 5, a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water vapor having a portion of a small amount of oxygen admixed therewith is brought into contact with the reforming catalyst and oxidizing catalyst mixture 52' and is thereby in part exothermally oxidized and reformed while rising its temperature. The gaseous mixture in part reformed is passed through the reforming reaction catalyst 51 and is thereby further reformed. Thereafter, the mixer 38 acts to admix a remainder of the small amount of oxygen into the further reformed gaseous mixture. Then, the resultant gaseous mixture past the reforming reaction catalyst and oxidizing catalyst mixture 52 raises its temperature to an aimed reaction termination temperature while being reformed by the action of the refining reaction catalyst, thus completing the reforming reaction.

[0055] Once the temperature rises occurs in it, the reforming reactor apparatus ceases feeding oxygen for admixing with the gaseous mixture.

[0056] The third form of embodiment of the invention described provides largely reducing the time required for starting the operation of the refining reactor apparatus

[0057] With the system so designed as have been described, it is made possible to reduce to minimum the total quantity of heat required for reformation and to reduce to an ultimate limit the amount of oxygen or air required for reformation while obtaining an increased concentration of hydrogen produced.

[0058] Although the present invention has been described hereinbefore in terms of the presently preferred forms of embodiments in respect of an auto-oxidation and internal heating type reforming method and apparatus for hydrogen production, it is to be understood that such disclosure is purely illustrative and is not to be interpreted as limiting. Consequently, without departing from the spirit and scope of the invention, various alterations, modifications, and/or alternative applications of the invention will, no doubt, be suggested to those skilled in the art after having read the preceding disclosure. Accordingly, it is intended that the following claims be interpreted as compassing all alterations, modifications, or alternative applications as fall within the true spirit and scope of the invention.

Claims

1. An auto-oxidation and internal heating type reforming method for hydrogen production for use in a process in which a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water vapor is fed into contact with a mass of a reforming catalyst to bring about a reforming reaction of the gaseous mixture therewith to produce hydrogen, characterized in that the method comprises the steps of:
 - admixing a small amount of an oxidizing catalyst with said reforming catalyst in said mass; and
 - admixing a small amount of oxygen with the gaseous mixture of said hydrocarbon or aliphatic alcohol with water vapor, whereby a portion of said hydrocarbon or aliphatic alcohol is exothermally oxidized to generate an amount of heat required to reform said gaseous mixture of said hydrocarbon or aliphatic alcohol with water vapor.
2. A reforming method set forth in claim 1, characterized in that said small amount of oxygen is admixed with the gaseous mixture of said hydrocarbon or aliphatic alcohol with water vapor prior to said gaseous mixture coming into contact with said reforming catalyst.
3. A reforming method as set forth in claim 1, characterized in that said small amount of oxygen is admixed with the gaseous mixture of said hydrocarbon or aliphatic alcohol with water vapor, after the gaseous mixture comes into contact with said reforming catalytic mass and is thereby in part reformed.
4. A reforming method set forth in claim 1, characterized in that said small amount of oxygen is in part admixed with the gaseous mixture of said hydrocarbon or aliphatic alcohol with water vapor prior to said gaseous mixture coming into contact with reforming catalyst, and a remainder of said small amount of oxygen is admixed with the resultant gaseous mixture after said resultant gaseous mixture comes in contact with said reforming catalyst and is thereby in part reformed.
5. A reforming method as set forth in any one of claims 1 to 4, characterized in that said oxidizing catalyst is distributively admixed with said reforming reaction catalyst in said mass.
6. A reforming method as set forth in any one of claims 1 to 5, characterized in that the method further comprises the step of passing air through a separation membrane to produce an oxygen gas having an increased concentration of oxygen which when exothermally reacted with said hydrocarbon or aliphatic alcohol is sufficient to generate a quantity of heat required for said reforming reaction.
7. A reforming method as set forth in any one of claims 1 to 6, characterized in that the method further comprises the steps of: producing a said small amount of oxygen from a water electrolytic cell; and feeding hydrogen generated from said water electrolytic cell into a region downstreams of a shift reaction.
8. A reforming method as set forth in any one of claims 1 to 7, characterized in that the method further comprises the steps of: using a reforming reaction catalyst loaded bed and a shift catalyst loaded bed partitioned by wall surfaces; and transferring heat generated by a shift reaction in said shift catalyst loaded bed into said reforming reaction catalyst loaded bed.
9. A reforming method as set forth in any one of claims 1 to 7, characterized in that the method further comprises the steps of: maintaining an operating pressure for said reforming reaction in a range between 6 kg/cm² and 8 kg/cm² and passing such gaseous mixtures including hydrogen through a membrane separator, thereby to raise a concentration of hydrogen while removing carbon monoxide (CO) generated as a by-product.
10. An auto-oxidation and internal heating type reforming reactor apparatus for hydrogen production for use in a process in which a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water

vapor is fed into contact with a mass of a reforming catalyst to bring about a reforming reaction of the gaseous mixture therewith to produce hydrogen, characterized in that the apparatus includes means for admixing a small amount of an oxidizing catalyst distributively with said reforming catalyst in said mass. 5

11. A reforming reactor apparatus as set forth in claim 10, characterized in that the apparatus further includes: a reforming reaction catalyst loaded bed having a reforming reaction catalyst and a small amount of an oxidizing catalyst distributively admixed therewith; and a thin layer of an oxidizing catalyst disposed behind and/or ahead of said reforming reaction catalyst loaded bed. 10 15

12. A reforming reactor apparatus as set forth in claim 10 or claim 11, characterized in that the apparatus includes a or said reforming reaction catalyst loaded bed and a shift catalyst loaded bed which are partitioned by wall surfaces and is configured so as to transfer heat generated by a shift reaction in said shift catalyst loaded bed to said reforming reaction catalyst loaded bed. 20 25

13. An auto-oxidation and internal heating type reforming reactor apparatus, characterized in that the apparatus includes an outer cylinder reaction catalyst bed loaded with a reforming reaction catalyst, and an inner cylinder catalyst bed comprising a reaction catalyst bed loaded with a or the reforming reaction catalyst containing a small amount of an oxidizing catalyst in an upstream side thereof and a shift catalyst bed loaded with a shift catalyst in a downstream side thereof, the apparatus being configured so as to transfer heat generated in said inner cylinder through wall surfaces thereof to said outer cylinder reaction catalyst bed, the apparatus further including a means for feeding a small amount of oxygen or air into a reaction mixture prior to said reaction mixture moving from an outer cylinder to an inner cylinder. 30 35 40 45 50 55

FIG. 1

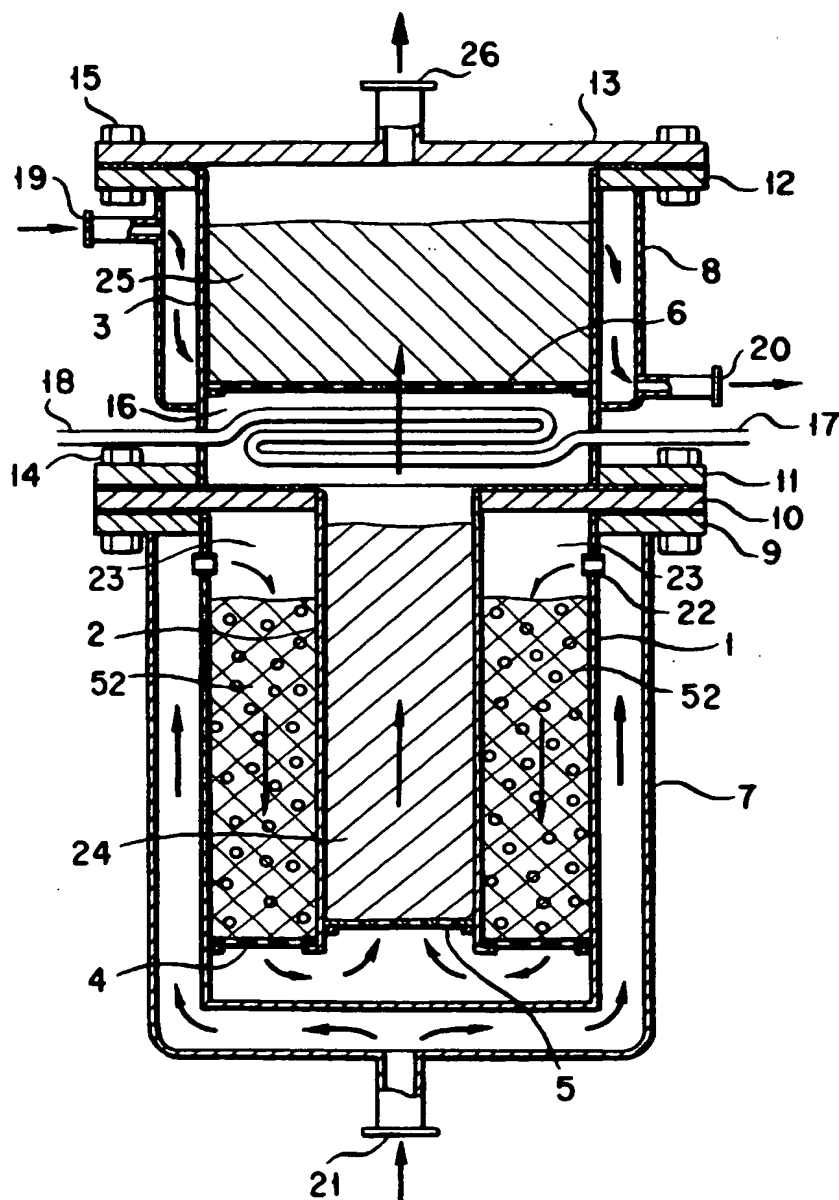


FIG. 2

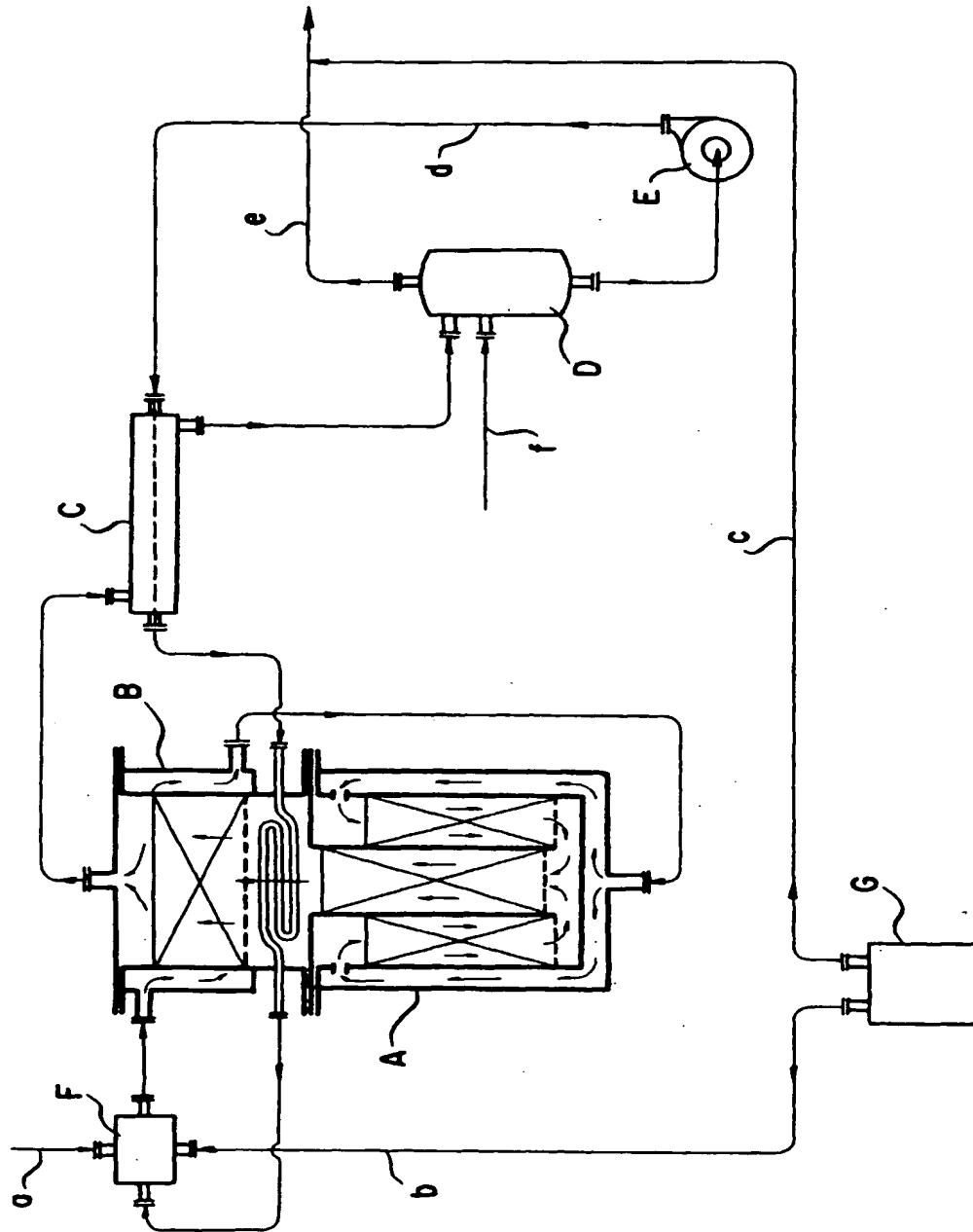


FIG. 3

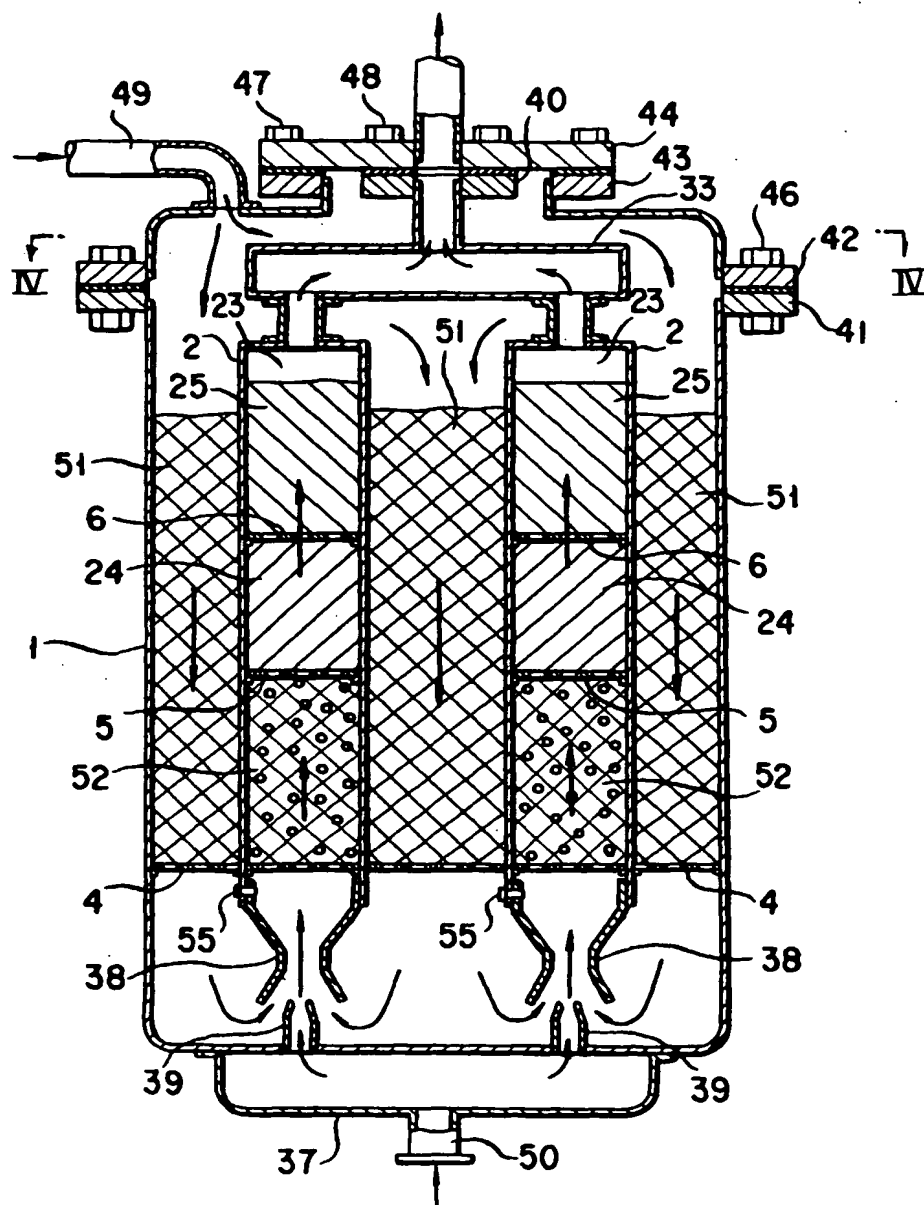


FIG. 4

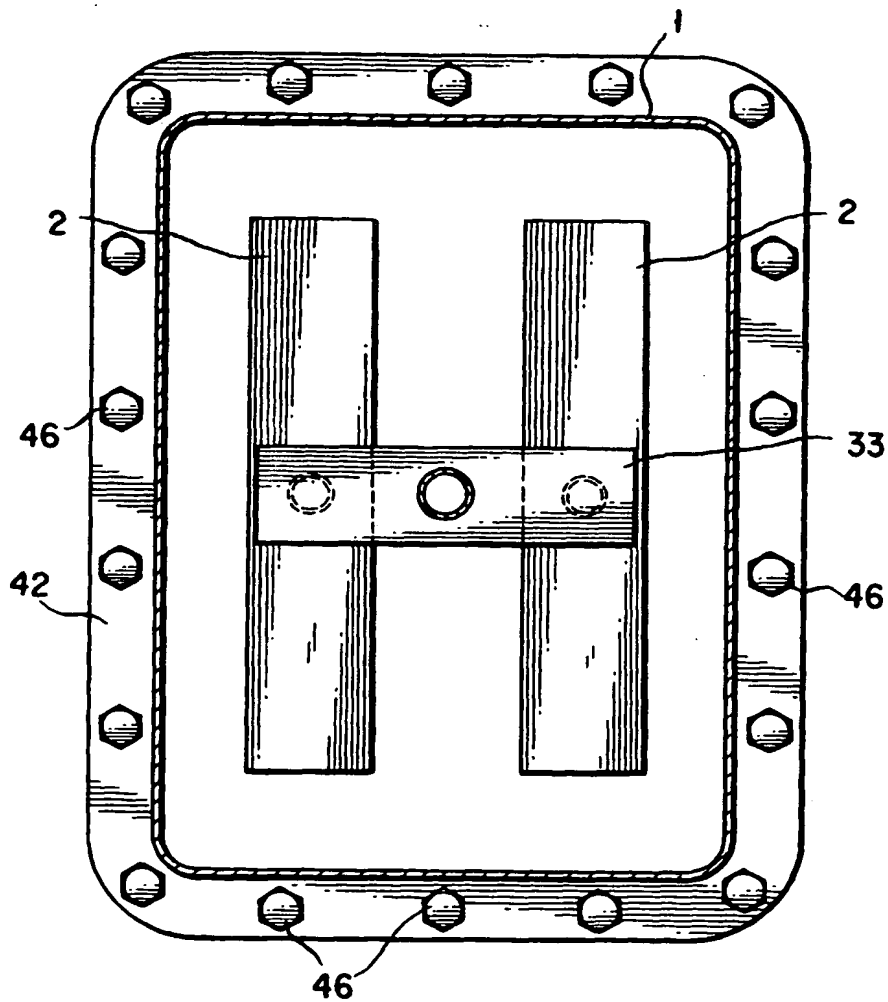
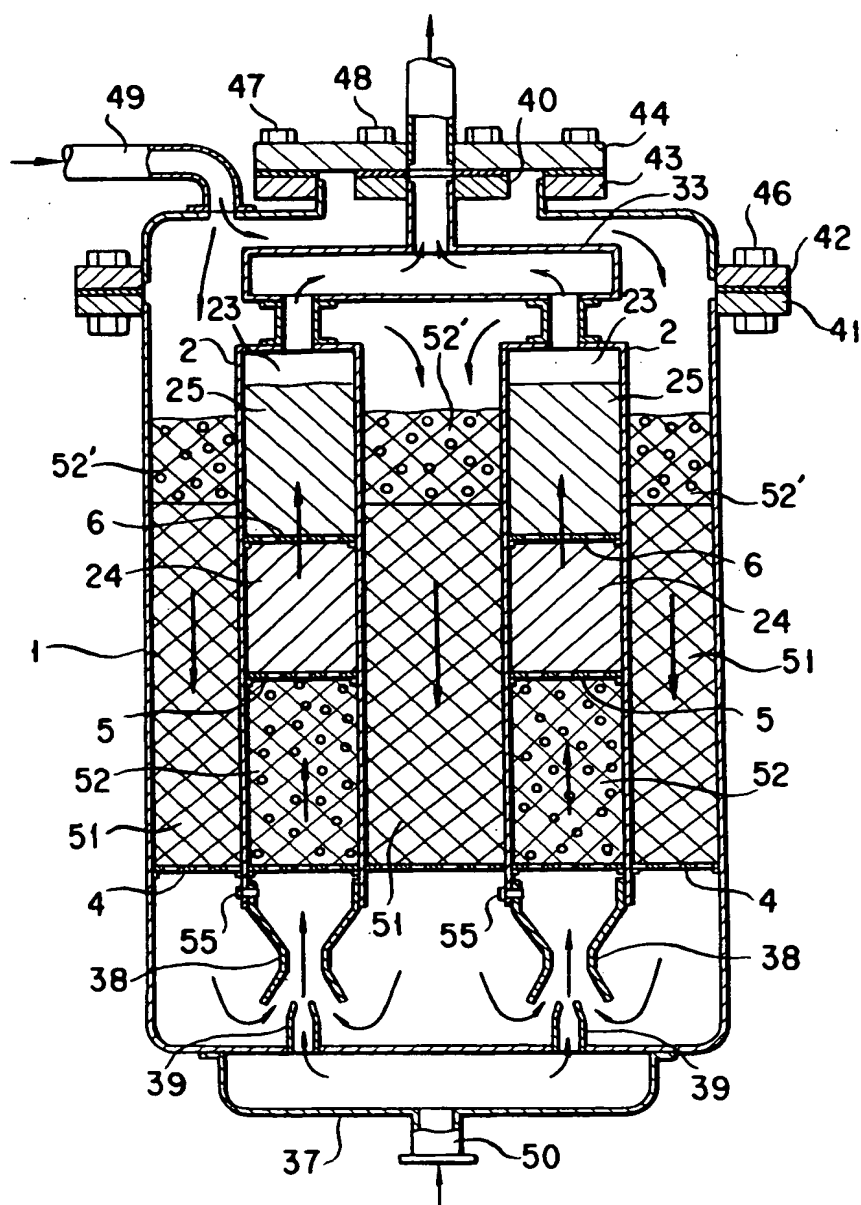


FIG. 5



(19)



Europäisches Patentamt
European Patent Office
Office européen des brevets



(11)

EP 1 094 030 A3

(12)

EUROPEAN PATENT APPLICATION

(88) Date of publication A3:
02.05.2002 Bulletin 2002/18

(43) Date of publication A2:
25.04.2001 Bulletin 2001/17

(21) Application number: **00122651.3**

(22) Date of filing: **18.10.2000**

(51) Int Cl.7: **C01B 3/38, C01B 3/32,
C01B 3/48, B01J 8/04,
C01B 3/40, C01B 3/50,
C01B 13/02**

(84) Designated Contracting States:
**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE**
Designated Extension States:
AL LT LV MK RO SI

(30) Priority: **20.10.1999 JP 29794599
03.03.2000 JP 2000058483**

(71) Applicant: **NIPPON CHEMICAL PLANT
CONSULTANT CO., LTD.
Chiyoda-ku Tokyo (JP)**

(72) Inventor: **Maruko, Saburo
Yamato-shi, Kanagawa (JP)**

(74) Representative: **Fuchs Mehler Weiss & Fritzsche
Patentanwälte
Postfach 46 60
65036 Wiesbaden (DE)**

(54) Reforming method and apparatus for hydrogen production

(57) An auto-oxidation and internal heating type reforming method and apparatus for hydrogen production are disclosed for use in a process in which a gaseous mixture of a hydrocarbon or an aliphatic alcohol with water vapor is fed (21) into contact with a mass of a reforming catalyst to bring about a reforming reaction of the gaseous mixture to produce hydrogen, wherein a small amount of an oxidizing catalyst is admixed with the re-

forming catalyst in that mass (52); and a small amount of oxygen is admixed with the gaseous mixture, whereby a portion of the hydrocarbon or aliphatic alcohol is exothermally oxidized to generate a quantity of heat required to reform the gaseous mixture of the hydrocarbon or aliphatic alcohol with water vapor.

EP 1 094 030 A3



European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 00 12 2651

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	WO 96 00186 A (JOHNSON MATTHEY PLC ; PIGNON JOHN FREDERICK (GB); EDWARDS NEIL (GB)) 4 January 1996 (1996-01-04) * claim 1; example 4 *	1, 2, 5	C01B3/38 C01B3/32 C01B3/48 B01J8/04 C01B3/40 C01B3/50 C01B13/02
X	US 5 112 578 A (MURAYAMA KATSUTOSHI ET AL) 12 May 1992 (1992-05-12) * claim 1; figure 1 *	1, 3, 5	
Y	IDEM	8, 10-12	
A	IDEM	13	
X	DE 44 23 587 A (DAIMLER BENZ AG) 11 January 1996 (1996-01-11) * claim 1; figures * * column 2, line 8 - line 22 * * column 2, line 53 - line 66 *	1, 2, 5, 9	
X	TAKI M ET AL: "HIGH EFFICIENT AND COMPACT METHANOL REFORMER FOR FUEL CELLS" TOYOTA TECHNICAL REVIEW, TOYOTA, JP, vol. 47, no. 2, November 1997 (1997-11), pages 76-81, XP002926587 ISSN: 0916-7501 * abstract *	1, 5	TECHNICAL FIELDS SEARCHED (Int.Cl.7) C01B B01J
X	WO 99 48805 A (CARPENTER IAN WILLIAM ; JOHNSON MATTHEY PLC (GB); HAYES JOHN WILLIA) 30 September 1999 (1999-09-30) * page 4, line 23 - line 26 * * claim 1; example 2 *	1, 2, 5	
-/--			
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 8 March 2002	Examiner Van der Poel, W
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document</p>			

EPO FORM 1503 03 82 (P04C01)



European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 00 12 2651

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	PATENT ABSTRACTS OF JAPAN vol. 012, no. 382 (C-535), 12 October 1988 (1988-10-12) & JP 63 129002 A (HITACHI LTD), 1 June 1988 (1988-06-01) * abstract *	1,2	
Y	IDEM	11	
Y	----- PATENT ABSTRACTS OF JAPAN vol. 1997, no. 03, 31 March 1997 (1997-03-31) & JP 08 301601 A (ISHIKAWAJIMA HARIMA HEAVY IND CO LTD), 19 November 1996 (1996-11-19) * abstract *	8,10,12	
A	----- WO 98 08771 A (LITTLE INC A) 5 March 1998 (1998-03-05) * figure 1 *	13	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 8 March 2002	Examiner Van der Poel, W
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document</p>			

EPO FORM 1503 03/92 (P4/C01)

**ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.**

EP 00 12 2651

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

08-03-2002

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
WO 9600186	A	04-01-1996	AT 174871 T	15-01-1999
			AU 683760 B2	20-11-1997
			AU 2749295 A	19-01-1996
			CA 2193115 A1	04-01-1996
			DE 69506869 D1	04-02-1999
			DE 69506869 T2	02-06-1999
			DK 766643 T3	23-08-1999
			EP 0766643 A1	09-04-1997
			ES 2126292 T3	16-03-1999
			WO 9600186 A1	04-01-1996
			JP 10502047 T	24-02-1998
			US 5762658 A	09-06-1998
US 5112578	A	12-05-1992	JP 1261201 A	18-10-1989
			AU 613175 B2	25-07-1991
			AU 3276789 A	19-10-1989
			CA 1328558 A1	19-04-1994
			DE 3912003 A1	26-10-1989
			GB 2222096 A , B	28-02-1990
DE 4423587	A	11-01-1996	DE 4423587 A1	11-01-1996
WO 9948805	A	30-09-1999	AU 3042399 A	18-10-1999
			EP 1068148 A1	17-01-2001
			WO 9948805 A1	30-09-1999
JP 63129002	A	01-06-1988	NONE	
JP 08301601	A	19-11-1996	NONE	
WO 9808771	A	05-03-1998	US 6126908 A	03-10-2000
			AT 203490 T	15-08-2001
			AU 729890 B2	15-02-2001
			AU 4161097 A	19-03-1998
			CN 1228749 A	15-09-1999
			DE 69705844 D1	30-08-2001
			EP 1118583 A2	25-07-2001
			EP 0922011 A2	16-06-1999
			ES 2159146 T3	16-09-2001
			JP 2000516902 T	19-12-2000
			WO 9808771 A2	05-03-1998
			US 6083425 A	04-07-2000
			US 6207122 B1	27-03-2001
			US 6123913 A	26-09-2000
			US 6254839 B1	03-07-2001
			US 2001009653 A1	26-07-2001

EPO FORM P0439

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82